metal-organic papers

Acta Crystallographica Section E Structure Reports Online

ISSN 1600-5368

Jens Hartung,^a* Arne Ludwig,^a Ingrid Svoboda^b and Hartmut Fuess^b

^aFachbereich Chemie, Organische Chemie, Technische Universität Kaiserslautern, Erwin-Schrödinger-Strasse, D-67663 Kaiserslautern, Germany, and ^bStrukturforschung, FB 11 Material- und Geowissenschaften, Technische Universität Darmstadt, Petersenstrasse 23, D-64287 Darmstadt, Germany

Correspondence e-mail: hartung@chemie.uni-kl.de

Key indicators

Single-crystal X-ray study T = 100 KMean σ (C–C) = 0.006 Å R factor = 0.055 wR factor = 0.160 Data-to-parameter ratio = 12.7

For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.

© 2007 International Union of Crystallography All rights reserved

{2,6-Dinitro-6-[(2-oxidophenyl)iminomethyl]phenolato}(ethanol)(ethanolato)oxidovanadium(V) at 100 K

The central V atom of the title complex, $[V(C_{13}H_7N_3O_6)-(C_2H_5O)O(C_2H_6O)]$, is positioned within a distorted octahedral coordination polyhedron formed by five O-donor atoms and one imine N atom. An axially bound ethanol solvent molecule forms an intermolecular hydrogen bond with the iminophenolate O atom of a neighboring complex, linking molecules into centrosymmetric dimers.

Comment

The title compound, (I), was prepared in order to study its reactivity toward dibasic tridentate Schiff base auxiliaries in ligand exchange reactions (Hartung *et al.*, 2007). It was crystallized from EtOH–pentane and investigated by X-ray diffraction at 100 K, in order to verify its structure.



In (I) (Fig. 1), the donor atoms form a distorted octahedral coordination polyhedron around the central V atom, which is displaced by 0.292 (2) Å toward O1, from the mean plane of atoms O2–O4/N1. A separation of 1.590 (3) Å indicates a double bond between V1 and O1. Longer distances are interpreted as single bonds connecting phenoxide O atoms O2 and O3 and alkoxide O atom O4 to V1 (Table 1; Rehder, 1999, Hartung *et al.*, 2004). The shortest of the V–O single bonds refers to alkoxide binding *trans* to the V1–N1 bond, which is a common structural motif in Schiff base-derived vanadium(V) complexes of the investigated type (Clague *et al.*, 1993, Bashirpoor *et al.*, 1997; Hartung *et al.*, 2004). The magnitude of the V1–N1 bond indicates a low affinitiy between the imine N atom and the central atom in neutral oxovanadium(V) complexes.

The H atom attached to O5 was located in a difference Fourier map restricting the data set to $2\theta < 40^{\circ}$. Refinement of the structure, using all data with the O–H distance fixed, provides an O5–H5O···H2A hydrogen bond that links pairs of molecules into centrosymmetric dimers. The chelate ligand in (I) is virtually planar, except for the nitro group connected to C9. This entity is tilted from the plane of the aryl substituent, presumably for reasons of steric and polar repulsion between atom O3 and atoms O8 and O9. Received 23 March 2007 Accepted 9 April 2007

Experimental

To a solution of 2-{[(2-hydroxy-3,5-dinitrophenyl)imino]methyl}phenol (303 mg, 1.00 mmol) in anhydrous EtOH (25 ml) was added VO(OEt)₃ (202 mg, 1.00 mmol) at 298 K. Stirring was continued for 24 h at 298 K. The solution was then concentrated under reduced pressure. The remaining residue was crystallized from EtOH-pentane (2:1 ν/ν) to afford crystals of (I) (m.p. > 623 K) suitable for X-ray diffraction. Analysis calculated for C₁₇H₁₈N₃O₉V: C 44.46, H 3.95, N 9.15%; found: C 43.88, H 3.76, N 9.15%; IR (KBr): 962 cm⁻¹ (V=O).

 $\beta = 106.38 \ (1)^{\circ}$

Z = 8

V = 3894.4 (8) Å³

Mo K α radiation $\mu = 0.57 \text{ mm}^{-1}$ T = 100 (2) K $0.44 \times 0.20 \times 0.04 \text{ mm}$

Diffraction, 2006)

 $R_{\rm int} = 0.040$

 $T_{\min} = 0.789, T_{\max} = 0.978$

10522 measured reflections

3781 independent reflections

2446 reflections with $I > 2\sigma(I)$

Crystal data

[V(C ₁₃ H ₇ N ₃ O ₆)(C ₂ H ₅ O)O-
$(C_2H_6O)]$
$M_r = 459.28$
Monoclinic, $C2/c$
$a = 21.109 (2) \text{\AA}$
b = 8.271 (1) Å
c = 23.249 (3) Å

Data collection

Oxford Diffraction Xcalibur diffractometer with Sapphire CCD detector Absorption correction: multi-scan (*CrysAlis RED*; Oxford

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.055$	H atoms treated by a mixture of
$wR(F^2) = 0.160$	independent and constrained
S = 1.05	refinement
3781 reflections	$\Delta \rho_{\rm max} = 1.12 \text{ e} \text{ \AA}^{-3}$
297 parameters	$\Delta \rho_{\rm min} = -0.56 \text{ e } \text{\AA}^{-3}$
1 restraint	

Table 1

Selected geometric parameters (Å, °).

914 (3)
755 (3)
293 (3)
)8 (12)
97 (12)
18 (13)
46 (11)
38 (11)
27 (11)
58 (11)
4.0 (6)
22

Table 2

Hydrogen-bond geometry (Å, °).

$D - H \cdots A$	D-H	$H \cdot \cdot \cdot A$	$D \cdots A$	$D - \mathbf{H} \cdot \cdot \cdot A$
$O5-H5A\cdots O2^{i}$	0.810 (10)	1.987 (14)	2.784 (4)	168 (4)
Symmetry code: (i) -	-x + 1, -y + 1, -	z.		

All H atoms, except for those attached to C14, C15, C16 and C17, were located in a difference Fourier map and refined giving C-H





Molecular structure of (I). Displacement ellipsoids are plotted at the 50% probability level.

distances in the range 0.78 (4)–1.12 (4) Å. Their displacement parameters were constrained with $U_{iso}(H)$ values set to $1.2U_{eq}$ of the parent atom. Atom H5A was refined with the O5–H distance restraint of 0.8 Å. H atoms attached to C14, C15, C16 and C17 were positioned geometrically and treated as riding atoms, with C–H distances in the range of 0.98–0.99 Å and with $U_{iso}(H)$ set to $1.2U_{eq}$ of the parent atom. The highest peak in the residual electron density map is located 1.56 Å from O1.

Data collection: *CrysAlis CCD* (Oxford Diffraction, 2006); cell refinement: *CrysAlis RED* (Oxford Diffraction, 2006); data reduction: *CrysAlis RED*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *PLATON* (Spek 2003) and *ORTEP-3* (Farrugia, 1997); software used to prepare material for publication: *SHELXL97*.

This work was supported by the Deutsche Forschungsgemeinschaft (grant No. Ha1705/8–2) and the Fonds der Chemischen Industrie.

References

- Bashirpoor, M., Schmidt, H., Schulzke, C. & Rehder, D. (1997). *Chem. Ber. Recl.* pp. 651–657.
- Clague, M. J., Keder, N. L. & Butler, A. (1993). Inorg. Chem. 32, 4754-4761.
- Farrugia, L. J. (1997). J. Appl. Cryst. 30, 565.
- Hartung, J., Demary, M., Ludwig, A. & Stapf, G. (2007). Vanadium the Versatile Metal, edited by D. Crans & K. Kustin, ACS Symposium Series. Washington, DC: American Chemical Society.
- Hartung, J., Drees, S., Greb, M., Schmidt, P., Svoboda, I., Fuess, H., Murso, A. & Stalke, D. (2004). Eur. J. Org. Chem. pp. 2388–2408.
- Oxford Diffraction (2006). CrysAlis CCD and CrysAlis RED. Versions 1.171.31.4. Oxford Diffraction, Abington, Oxfordshire, England.
- Rehder, D. (1999). Coord. Chem. Rev. 182, 297–322.
- Sheldrick, G. M. (1997). SHELXS97 and SHELXL97. University of Göttingen, Germany.
- Spek, A. L. (2003). J. Appl. Cryst. 36, 7-13.